Pushing the limits of energetic materials – the synthesis and characterization of dihydroxylammonium 5,5′-bistetrazole-1,1′-diolate†

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The safe preparation and characterization (XRD, NMR and vibrational spectroscopy, DSC, mass spectrometry, sensitivities) of a new explosive dihydroxylammonium 5,5′-bistetrazole-1,1′-diolate (TKX-50) that outperforms all other commonly used explosive materials is detailed. While much publicized high-performing explosives, such as octanitrocubane and CL-20, have been at the forefront of public awareness, this compound differs in that it is simple and cheap to prepare from commonly available chemicals. TKX-50 expands upon the newly exploited field of tetrazole oxide chemistry to produce a material that not only is easily prepared and exceedingly powerful, but also possesses the required thermal insensitivity, low toxicity, and safety of handling to replace the most commonly used military explosive, RDX (1,3,5-trinitro-1,3,5-triazacyclohexane). In addition, the crystal structures of the intermediates 5,5′-bistetrazole-1,1′-diol dihydrate, 5,5′-bistetrazole-1,1′-diol dimethanolate and dimethylammonium 5,5′-bistetrazole-1,1′-diolate were determined and presented.

Introduction

The rational design of new energetic materials is a rapidly exploding field1–7 with a long traditional rooting in the chemical sciences8,9 and a complexity that rivals that of the drug design. While the field has come a long way since the days of Liebig, Berzelius and Gay-Lussac, and the concept of isomerism being determined from explosive silver fulminate and non-explosive silver cyanate,10 current work in this field still follows the trend of its historic beginnings; that of simultaneous academic and practical interest and advances. In the quest for higher-performing, safer, cheaper, greener, explosive materials, energetic materials chemistry must push the boundaries of the energy capacity of compounds, which requires new classes of compounds,2–11 new synthetic strategies,12 and advanced computational techniques. For example, the high nitrogen content of many advanced explosives has led to the preparation of new nitrogen–nitrogen bond forming reactions12 and new heterocyclic systems13 in the quest for even higher performance.

In both civilian and military circles, the highest performing explosives make use of the same strategy: cyclic and caged nitramines. Belonging to the oldest class of explosives, those derive their energy from the oxidation of a carbon backbone by containing the oxidizer in the same molecule; RDX (1,3,5-trinitro-1,3,5-triazacyclohexane), HMX (1,3,5,7-tetranitro-1,3,5,7-tetraazacyclotetane) and CL-20 (2,4,6,8,10,12-hexanitro-2,4,6,8,10,12-hexaza-isowurtzitane) all have fatal flaws that mandate replacement with modern explosives. Advanced energetic strategies allow for retention or improvement of the explosive performance, while avoiding the multitude of downsides present in these compounds: toxicity to living organisms (all), difficult and expensive synthesis (HMX, CL-20), high sensitivity to mechanical stimuli (all), and spontaneous changing of properties (CL-20).2 New strategies in the design of energetic materials include those with ring or cagecontaining strong dipoles or zwitterionic structures.14 Unfortunately, the known materials with the highest detonation energy are often highly sensitive due to their unprecedented energy content,5 and are made via long and expensive pathways with a multitude of steps, making industrial scale-up infeasible. For example, both DDF (dinitroazofuroxane) and ONC (octanitrocubane) possess detonation velocities at the limit of known performances (around 10 000 m s−1), however both are highly

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† Electronic supplementary information (ESI) available: 1. Materials and methods; 1.1. NMR spectroscopy; 1.2. vibrational spectroscopy; 1.3. mass spectrometry and elemental analysis; 1.4. differential scanning calorimetry; 1.5. sensitivity testing; 2. experimental work; 2.1. synthesis via oxidation of 5,5′-bistetrazole with potassium peroxyxmonosulfate; 2.2. synthesis via cyclization of diazidoglyoxime; 2.3. safer synthesis including a multi-step one pot reaction; 3. X-ray diffraction; 3.1. instrument and refinement software; 3.2. crystallographic data and refinement parameters; 3.3. bond lengths, bond angles and hydrogen bonding of TKX-50; 3.4 crystal structures of 5,5′-bistetrazole-1,1′-diol; 3.5. crystal structure of dimethylammonium 5,5′-bistetrazole-1,1′-diolate; 4. explosive performance; 4.1. heat of formation calculations; 4.2. small scale shock reactivity test; 4.3 flame test; 4.4 hot plate test; 5. toxicity assessment; 6. Fast Cook-Off test, cif files; CCDC 872230, 872231, 872232, 884559, 884560 and 884561. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c2jm33646d
sensitive and have more than 10 synthetic steps with exotic, expensive reagents used.\textsuperscript{15}

A very promising explosphoric moiety in the design of new energetic materials is the tetrazole ring; the carbon on position 5 of the ring allows the facile attachment of various substituents for energetic tailorability, and the high nitrogen content and heat of formation of the heterocycle lead to high energetic performances. In order to improve the energetic properties of tetrazoles, several recently published studies showed that introduction of N-oxides yields compounds with even higher densities and stabilities, lower sensitivities and better oxygen balances.\textsuperscript{5,7,11} Combining these principles with practical considerations in mind, a simple and secure synthetic pathway to the high performing energetic material dihydroxylammonium 5,5'-bistetrazole-1,1'-diolate (TKX-50) was devised.

**Results and discussion**

**Synthesis (simple and scalable)**

There are two major routes (A and B) to the 5,5'-bistetrazole-1,1'-diol (1,1-BTO) moiety (Scheme 1). The first (A) of which, the oxidation of the parent heterocycle with aqueous potassium peroxymonosulfate only leads to 1,1-BTO in poor yield (11%). The oxidation of the 5,5'-bistetrazolate anion with peroxymonosulfate was carried out in a manner similar to that we have previously reported for 5-nitro- and 5-azidotetrazoles.\textsuperscript{5} Unfortunately, this reaction was found to produce the 2,2' isomer as the major product, with only traces of the 1,1' isomer which crystallized upon adding aqueous hydroxylamine.

After discovering the outstanding characteristics of TKX-50 as a highly explosive, a different route to the precursor 5,5'-bistetrazole-1,1'-diol was necessitated. Tselinskii et al.\textsuperscript{16} reported on the synthesis of the mentioned precursor 1,1-BTO from the cyclization of diazidoglyoxime under acidic conditions for the first time. Diazidoglyoxime is prepared from dichloroglyoxime in a chloro–azido exchange reaction in DMF with more than 80 % yield, whereas dichloroglyoxime is prepared from glyoxime \textit{via} chlorination in ethanol in high yield.

The problematic step here is the isolation of the highly friction and impact sensitive compound diazidoglyoxime, mandating a revised procedure before industrial-scaled-up. The problem was overcome by a procedure combining the formation and cyclization of diazidoglyoxime in one step in solution. Starting from commercially available glyoxal, the reaction process was transformed into a five step, four pot synthesis to isolate TKX-50.

The prepared solution of diazidoglyoxime in DMF (impure with sodium chloride) is directly poured into diethylether and 1,1'-dichloroethane. The diethyl ether is allowed to evaporate at room temperature and the mixture is left for 24 hours. The obtained yellowish mass is washed with chloroform, ethanol and methanol and then dried in vacuo. The pure compound is recrystallized from water or ethanol.

**Scheme 1** Synthesis of TKX-50 \textit{via} oxidation of 5,5'-bistetrazole (A) and \textit{via} cyclization of diazidoglyoxime (B).

HCl gas is bubbled through (Scheme 2). After cyclization of the azidooxime in the acidic medium the dimethylammonium salt of 5,5'-bistetrazole-1,1'-diol is formed by a reaction with dimethylamine (formed by hydrolysis of DMF). After isolation and recrystallization of dimethylammonium 5,5'-bistetrazole-1,1'-diolate, it is dissolved in a sufficient amount of boiling water and combined with a solution of hydroxylammonium chloride, from which TKX-50 crystallizes first (Scheme 3).

An alternative procedure using NMP (\textit{N}-methyl-2-pyrrolidone) instead of DMF for the chloro–azido exchange, followed by the same treatment, leads to the free acid 5,5'-bistetrazole-1,1'-diol which is then isolated as its sodium salt tetrahydrate upon the addition of aqueous sodium hydroxide and subsequently treated with hydroxylammonium chloride in water. Starting from dichloroglyoxime, the overall yields of both procedures are very high with 72 % (DMF-route) and 85 % (NMP-route) for the synthesis of TKX-50. For a detailed description of all synthetic routes yielding TKX-50 and for all analytical data please refer to the ESI.\textsuperscript{†}

**X-ray diffraction**

The crystal structure of TKX-50 was determined at three temperatures (100 K, 173 K, 298 K) in order to detect potential low temperature phase transitions and obtain precise densities (for explosive performance calculations). In addition the crystal structures of the intermediates 5,5'-bistetrazole-1,1'-diol dihydrate (recryst. from either water, MeCN, EtOH or glacial acetic acid), 5,5'-bistetrazole-1,1'-diol dimethanolate (recryst. from methanol) and dimethylammonium 5,5'-bistetrazole-1,1'-diolate (crystallized from H\textsubscript{2}O) were determined and are presented in the ESI.\textsuperscript{†} Detailed crystallographic data and parameters of the measurements and solutions are given in Table S1.\textsuperscript{†} The lack of observed phase transitions between 100 K and 298 K is advantageous for energetic materials use as constant properties upon temperature changes result. The density follows the expected trend of decreasing with increased temperature (100 K: 1.918 g cm\textsuperscript{-3} > 173 K: 1.915 g cm\textsuperscript{-3} > 298 K: 1.877 g cm\textsuperscript{-3}). TKX-50 crystallizes in the monoclinic space group \textit{P}2\textsubscript{1}/\textit{c} with two anion–cation moieties in the unit cell. The molecular moiety of...
Fig. 1 Representation of the solid state molecular structure of TKX-50 at 100 K. Thermal ellipsoids are drawn at the 50% probability level; symmetry codes: (i) 2 – x, –y, 2 – z; (ii) x, 0.5 – y, –0.5 + z; and (iii) –1 + x, 0.5 – y, –0.5 + z.

TKX-50 at 100 K is depicted in Fig. S1† Its density of 1.918 g cm\(^{-3}\) is significantly higher than that of non-oxide dihydroxyl-ammonium 5,5'-bistetrazolate (1.742 g cm\(^{-3}\)) recently published.\(^{17}\) The reason for this may be the strong hydrogen bond on the chemical equilibrium, a steady state model of detonation.

Energetic performance

The performance data (Table 1) were calculated with the computer code EXPLO5.05 (latest version). EXPLO5.05 is based on the chemical equilibrium, a steady state model of detonation. It uses Becker–Kistiakowsky–Wilson’s equation of state (BKW EOS) for gaseous explosive mixture and Cowan–Fickett’s equation of state for solid carbon (see ESI†). The input is based on the sum formula, calculated heats of formation (see ESI†) and the maximum densities according to their crystal structures (ESI, Table S1†).

With respect to the detonation velocity (Table 1), TKX-50 shows higher calculated values than all other mass-produced and used explosives like 2,4,6-trinitrotoluene (TNT), RDX, HMX and CL-20. Looking at the detonation pressure, TKX-50 exceeds the values calculated for TNT and RDX and is comparable to HMX, but is slightly lower than for CL-20. Also in terms of potential use as a propellant mixture ingredient TKX-50 shows promising values due to its high nitrogen content. The calculated specific impulse using 60 bar rocket conditions is 261 seconds, which is slightly better than those of the other compounds in Table 1.

To assess the explosive performance of TKX-50 on a small laboratory scale, a small-scale reactivity test (SSRT) was carried out (see ESI†) in comparison to CL-20 and RDX. Here, a defined volume of the explosive is pressed into a perforated steel block, which is topped with a commercially available detonator (Orica, DYNADETEC-C2-0ms). Initiation of the tested explosive results in denting a separate aluminium block, which is placed right underneath the steel block (Fig. 2). From measuring the volumes of the dents (CL-20 > TKX-50 > RDX) (Table 2 in the ESI†), it can be concluded that the small scale explosive performance of TKX-50 exceeds the performance of commonly used RDX and is comparable to that of CL-20.

The performance and safety characteristics for shipping of an explosive can be related to the data obtained from the Koenen test.\(^{18,19}\) The explosive is placed in an open-ended, flanged steel tube, which is locked up with a closing plate with variable orifice (0–10 mm), through which gaseous decomposition products are passed and the sand and the steel tube are observed. The explosive is classified as a black powder (CL-20) or as a white powder (TKX-50 and RDX). Under the Koenen test, the explosive is considered safe for transportation if no propagation of the reaction is observed. Under the Koenen test, TKX-50 exceeds the performance of CL-20 and RDX and is comparable to HMX.

Table 1 Energetic properties and detonation parameters of prominent high explosives in comparison to TKX-50

<table>
<thead>
<tr>
<th></th>
<th>2,4,6-TNT</th>
<th>RDX</th>
<th>β-HMX</th>
<th>ε-CL-20</th>
<th>TKX-50</th>
</tr>
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<tbody>
<tr>
<td>Formula</td>
<td>C(_6)H(_3)N(_2)O(_6)</td>
<td>C(_6)H(_4)N(_2)O(_6)</td>
<td>C(_4)H(_8)N(_3)O(_5)</td>
<td>C(_4)H(_6)N(_2)O(_12)</td>
<td>C(_4)H(_8)N(_2)O(_4)</td>
</tr>
<tr>
<td>Molecular mass [g mol(^{-1})]</td>
<td>227.13</td>
<td>222.12</td>
<td>296.16</td>
<td>438.19</td>
<td>236.15</td>
</tr>
<tr>
<td>IS [J](^{b})</td>
<td>15(^{21})</td>
<td>7.5(^{21})</td>
<td>7(^{21})</td>
<td>4 (^{21})</td>
<td>20</td>
</tr>
<tr>
<td>FS [N]f</td>
<td>353</td>
<td>120(^{21})</td>
<td>112(^{21})</td>
<td>48 (^{21})</td>
<td>120</td>
</tr>
<tr>
<td>ESD-test [J](^{f})</td>
<td>—</td>
<td>0.20</td>
<td>0.20</td>
<td>0.13</td>
<td>0.10</td>
</tr>
<tr>
<td>Ω [%]f</td>
<td>18.50</td>
<td>37.84</td>
<td>37.84</td>
<td>38.3</td>
<td>59.3</td>
</tr>
<tr>
<td>T(_m) [°C]g</td>
<td>81</td>
<td>205(^{22})</td>
<td>275 (^{24})</td>
<td>—</td>
<td>—</td>
</tr>
<tr>
<td>T(_{det}) [°C]g</td>
<td>290</td>
<td>210(^{24})</td>
<td>279 (^{24})</td>
<td>215 (^{23})</td>
<td>221</td>
</tr>
<tr>
<td>Density [g cm(^{-3})]</td>
<td>1.713 (100 K)(^{24})</td>
<td>1.858 (90 K)(^{26})</td>
<td>1.944 (100 K)(^{28})</td>
<td>2.083 (100 K)(^{29})</td>
<td>1.918 (100 K)(^{30})</td>
</tr>
<tr>
<td>Theor. Δ(H) [kJ mol(^{-1})]</td>
<td>1680</td>
<td>86.3</td>
<td>116.1</td>
<td>365.4</td>
<td>446.6</td>
</tr>
<tr>
<td>Theor. Δ(U) [kJ kg(^{-1})]</td>
<td>—</td>
<td>1680</td>
<td>492.5</td>
<td>918.7</td>
<td>2006.4</td>
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<tr>
<td>EXPLO5.05 values</td>
<td>—</td>
<td>—</td>
<td>—</td>
<td>—</td>
<td>—</td>
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<tr>
<td>Δ(U) [kJ kg(^{-1})]</td>
<td>5258</td>
<td>6190</td>
<td>6185</td>
<td>6406</td>
<td>6025</td>
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<tr>
<td>T(_b) [K]h</td>
<td>3563</td>
<td>4232</td>
<td>4185</td>
<td>4616</td>
<td>3954</td>
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<tr>
<td>P(_c) [kbar]i</td>
<td>235</td>
<td>380</td>
<td>415</td>
<td>467</td>
<td>424</td>
</tr>
<tr>
<td>D [m s(^{-1})]j</td>
<td>7459</td>
<td>8983</td>
<td>9221</td>
<td>9455</td>
<td>9698</td>
</tr>
<tr>
<td>Gas vol. [L kg(^{-1})]k</td>
<td>569</td>
<td>734</td>
<td>729</td>
<td>666</td>
<td>846</td>
</tr>
<tr>
<td>(I)_(_p) [s]k</td>
<td>205</td>
<td>258</td>
<td>258</td>
<td>251</td>
<td>261</td>
</tr>
</tbody>
</table>

\(^a\) Impact sensitivity (BAM drophammer (1 of 6)). \(^b\) Friction sensitivity (BAM friction tester (1 of 6)). \(^c\) Electrostatic discharge device (OZM research).

\(^d\) Nitrogen content. \(^e\) Oxygen balance (Ω = (xO – 2yC – 1/2zHM)/1600). \(^f\) Decomposition temperature from DSC (β = 5 °C min\(^{-1}\)). \(^g\) From X-ray diffraction. \(^h\) Calculated (CBS-4M method) enthalpy of formation. \(^i\) Calculated energy of formation. \(^j\) Energy of explosion. \(^k\) Explosion temperature. \(^l\) Detonation pressure. \(^m\) Detonation velocity. \(^n\) Volume of detonation gases (assuming only gaseous products). \(^o\) Specific impulse using isobaric (60 bar) conditions.
vented. A defined volume of 25 mL of the compound is loaded into the flanged steel tube and a threaded collar is slipped onto the tube from below. The closing plate is fitted over the flanged tube and secured with a nut. The decomposition is initiated via thermal ignition using four Bunsen burners, which are ignited simultaneously. The test is completed when either rupture of the tube or no reaction is observed after heating the tube for a minimal time period of at least 5 min. In the case of the tube’s rupture, the fragments are collected and weighed. The reaction is evaluated as an explosion if the tube is destroyed into three or more pieces. The Koenen test was performed with 23.0 g of TKX-50 using a closing plate with an orifice of 10 mm and caused the rupture of the steel tube into approximately 100 pieces, the size of which reached down to smaller than 1 mm from 40 mm (Fig. 3).

TNT destroys the steel tube up to an orifice width of 6 mm, RDX even up to 8 mm. In order to get an “Interim Hazard Classification” also a “Fast Cook-Off test” (UN test 3d) was performed in which TKX-50 underwent controlled deflagration (no explosion occurred).

High safety – low sensitivity

Impact sensitivity is a high priority in explosive devices used in the military due to the range of stresses devices may be exposed to. The impact sensitivity of TKX-50 is 20 J which is much lower than those for RDX, HMX and CL-20, which range from 4 to 7.5 J, and all three of which need desensitizing components added for practical use. The low impact sensitivity of TKX-50 shows that it can be used without desensitization.

Friction sensitivity is more important in the manufacturing context, where TKX-50 with 120 N is of comparable or lower sensitivity than any of RDX, HMX or CL-20, increasing the margin of safety in the industrial context. Both the impact and friction sensitivities of TKX-50 as compared to 2,4,6-trinitrotoluene (TNT), RDX, HMX and CL-20 are presented in Table 1.

The human body can generate up to 25 mJ of static electricity, which can easily set off the most sensitive explosives such as lead azide or silver fulminate. TKX-50 has an electrostatic sensitivity of 0.100 J, which is far higher than the human body can generate, allowing a comparable margin of safety when handling, comparable to RDX or HMX.

Thermal stability is important for any explosive in practical use as demanding military requirements need explosives that can withstand high temperatures. For example, a munition sitting in the desert can exceed 100 °C and for general use a component explosive must be stable above 200 °C. TKX-50 with a decomposition onset of 222 °C easily surpasses this requirement (Fig. 4, inset). This stability has been confirmed using a long-term stability test, where the sample is heated in an open glass vessel to a temperature of 75 °C over 48 h to ensure safe handling of the material even at elevated temperatures (Fig. 4, outer curve). The lack of exothermic or endothermic events in the sample temperature or heat flow curve implies that the compound is stable.

Toxicity – environmentally friendly

One of the major aims in our search for new “green” energetic materials is the low toxicity of the newly investigated compound
Fig. 5 Toxicity assessment of TKX-50 using a luminescent bacteria inhibition test. Plot of log $T$ against log $c$ for determination of the EC$_{50}$ value.

itself, and of its degradation and decomposition products. In recent times the toxicity of energetic materials is a growing concern due to new understandings of the fate of explosives in the environment. The nitramine content of the ubiquitous RDX as well as less used HMX and CL-20 has been shown to be toxic to vital organisms at the base of the food chain, and in addition RDX is a probable human carcinogen. To assess the toxicity of TKX-50 to aquatic life, diluted aqueous solutions of the high explosive were subjected to the luminescent marine bacterium *Vibrio fischeri* using the commercially available bioassay system LUMISTox®. *Vibrio fischeri* is a representative species for other aquatic life and therefore a useful indicator when it comes to groundwater pollution. Being the most important toxicological parameter, the EC$_{50}$ value of the sample was determined. EC$_{50}$ is the effective concentration of the examined compound, at which the bioluminescence of the strain *Vibrio fischeri* is decreased by 50% after a defined period of exposure as compared to the original bioluminescence of the sample before being treated with the differently diluted solutions of the test compound. For RDX we observe an EC$_{50}$ value of 91 ppm after an incubation time of 30 minutes. The herein determined EC$_{50}$ value of TKX-50 of 130 ppm (Fig. 5 and ES†) lies significantly above the EC$_{50}$ value found for RDX indicating a lower toxicity to *Vibrio fischeri*, and as such, other aquatic life.

Conclusions

We have detailed the preparation of a new explosive, TKX-50 or dihydroxylammonium 5,5'-bistetrazole-1,1'-diolate. This material has exemplified the utility of the tetratetra N-oxide chemistry by providing a new explosive material that is of very high performance (as calculated and demonstrated by SSRT testing), pushing the limits towards the most powerful explosives known, and synthesized in an industrially viable process. Additionally, TKX-50 is of lower sensitivity (mechanically and thermally) than its contemporaries in currently used explosives such as RDX, HMX and CL-20, making increased margins of safety when applied in practical use and devices. Finally, we have demonstrated the lower toxicity of TKX-50 compared to the nitramine RDX, as determined by the EC$_{50}$ value for the decrease in luminescence of *Vibrio fischeri*. All of the characteristics of TKX-50 make it appropriate and exemplary to not just fulfill the long-standing goal of a “green” RDX replacement, but also to replace it with a material of superior performance.

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Notes and references

30 X-ray density (for details see CCDC 872231 and 872232).